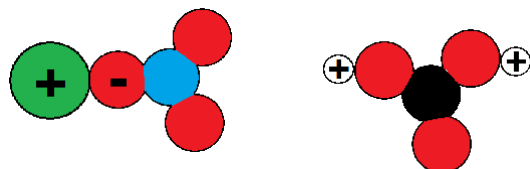
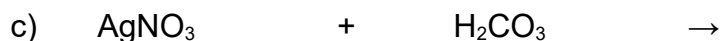
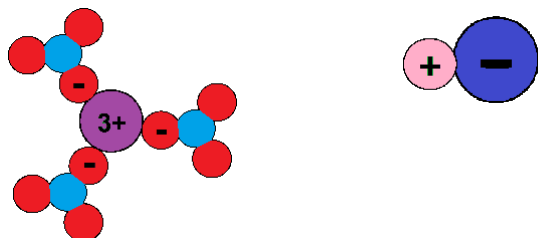
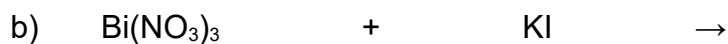
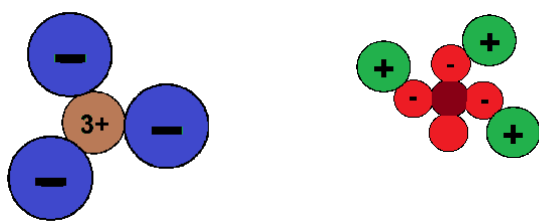


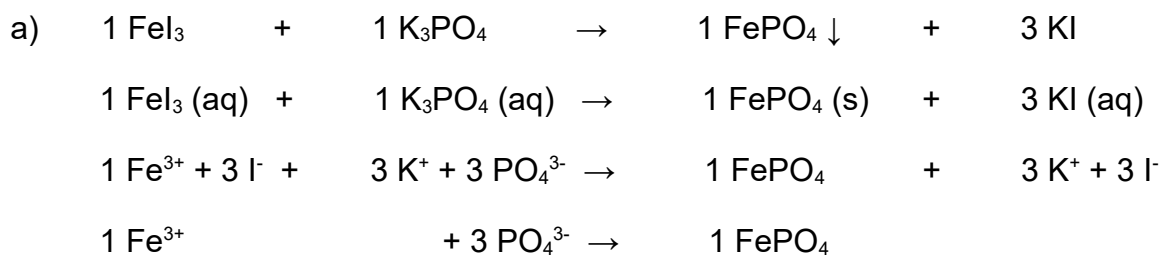
Activité précipitation

N'hésitez pas à faire / refaire pour entraînement les différentes réactions de précipitation proposées au cours 19 et à visualiser les modèles moléculaires laissés par vos camarades pour vérifier votre bonne compréhension du modèle microscopique qui explique les observations macroscopiques !

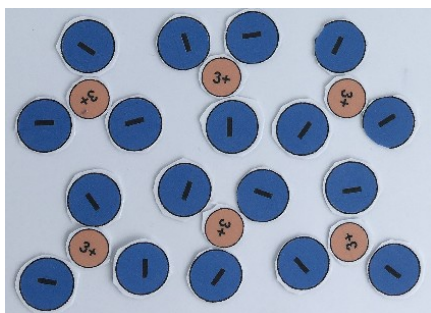
Pour chacune des réactions, on vous demandait : les équations globales, ioniques et ioniques réduites ainsi que les représentations moléculaires des réactifs à l'état solide, des réactifs après dissociation dans l'eau et des produits de la réaction.

Les quatre réactions proposées, dont le corrigé suit, étaient les suivantes :

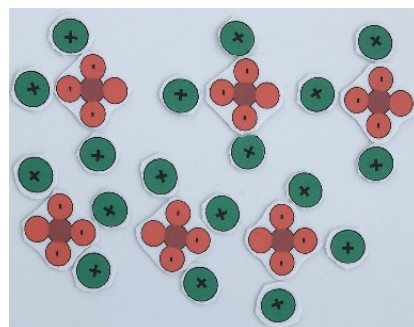




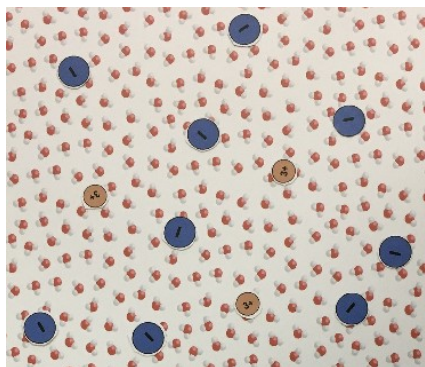
$\text{FeI}_3 (\text{s}) :$



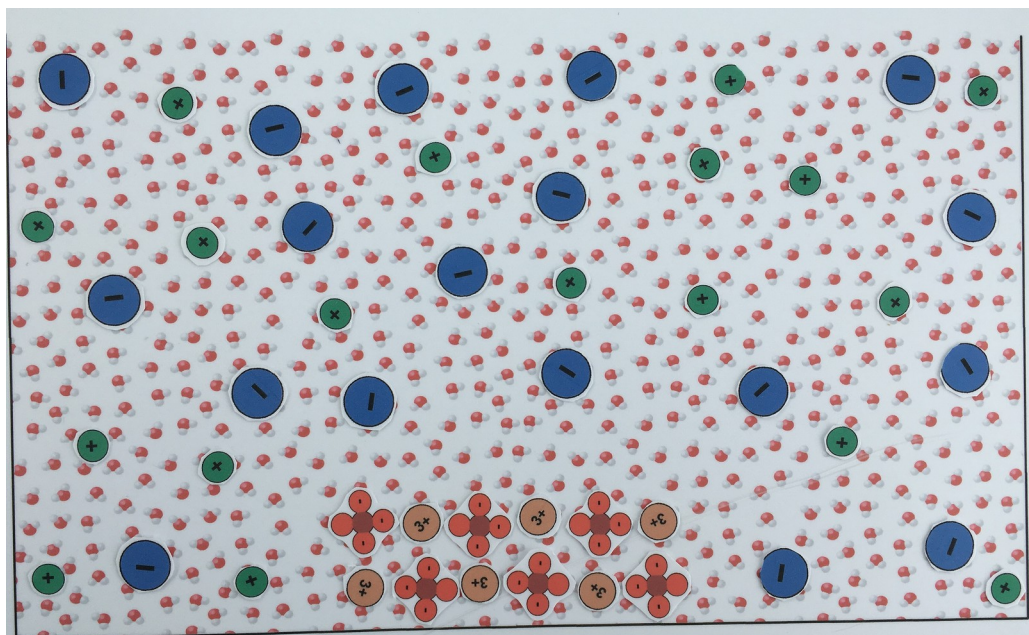
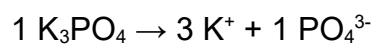
$\text{K}_3\text{PO}_4 (\text{s}) :$

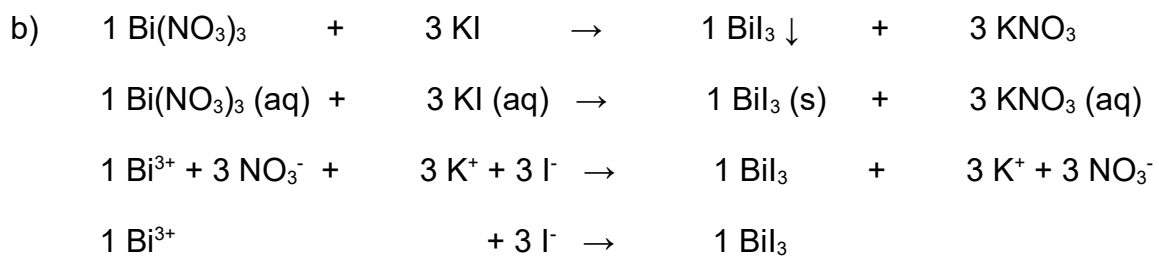


$\text{FeI}_3 (\text{aq}) :$

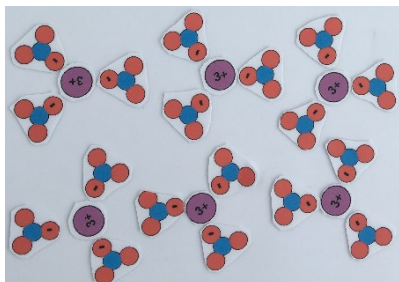


$\text{K}_3\text{PO}_4 (\text{aq}) :$

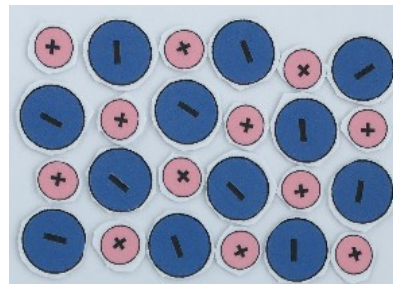




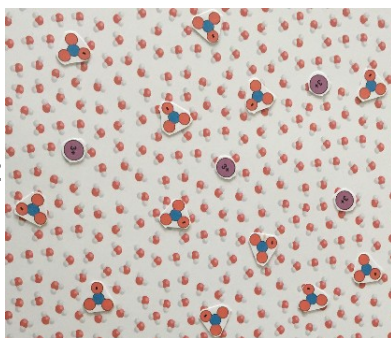
$\text{Bi(NO}_3)_3 (\text{s}) :$



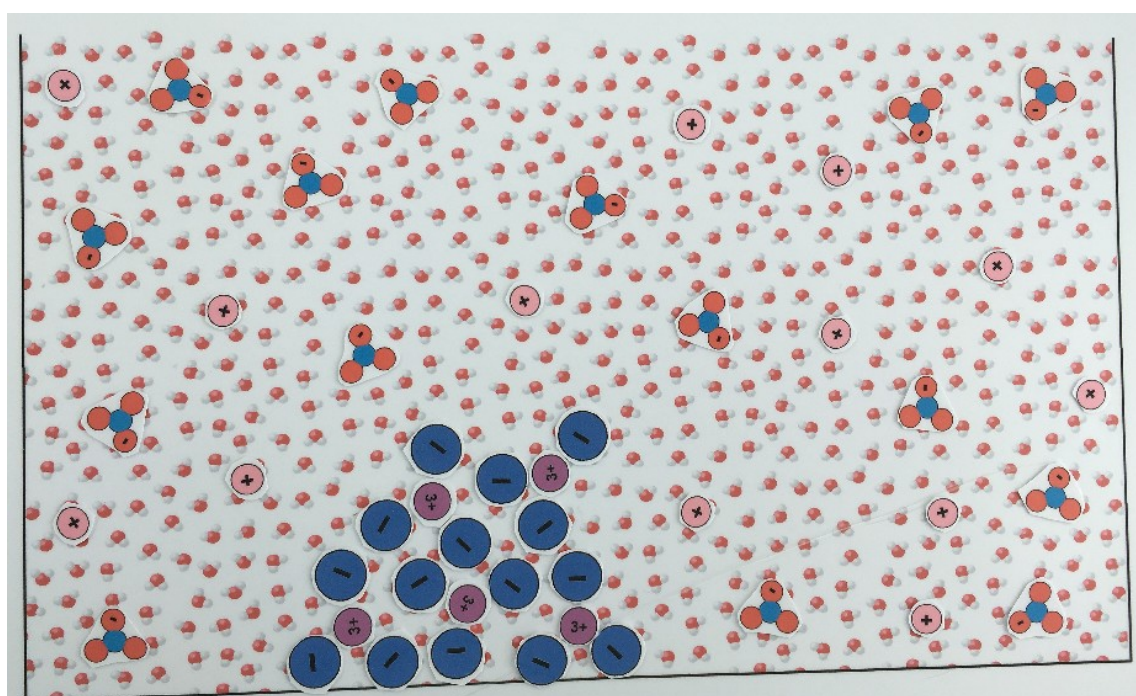
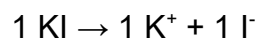
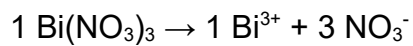
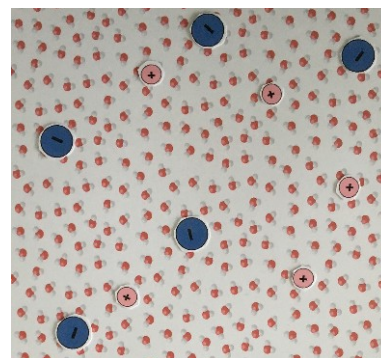
$\text{KI} (\text{s}) :$

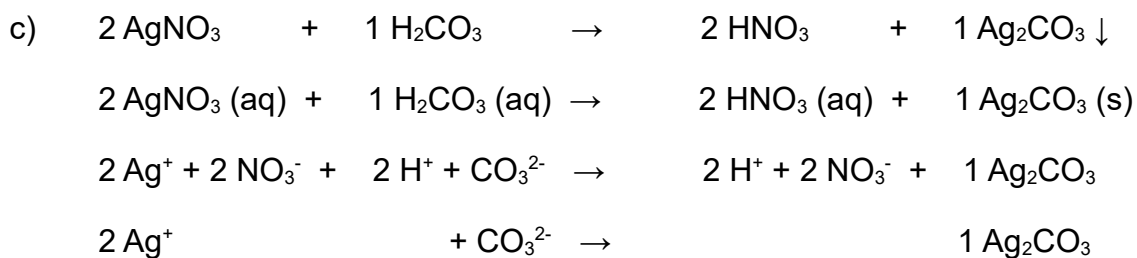


$\text{Bi(NO}_3)_3 (\text{aq}) :$

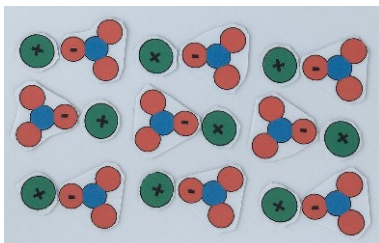


$\text{KI} (\text{aq}) :$

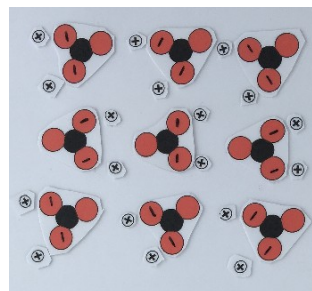




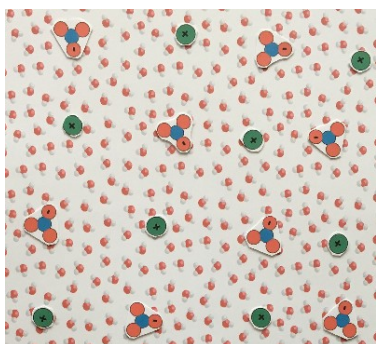
$\text{AgNO}_3 (\text{s}) :$



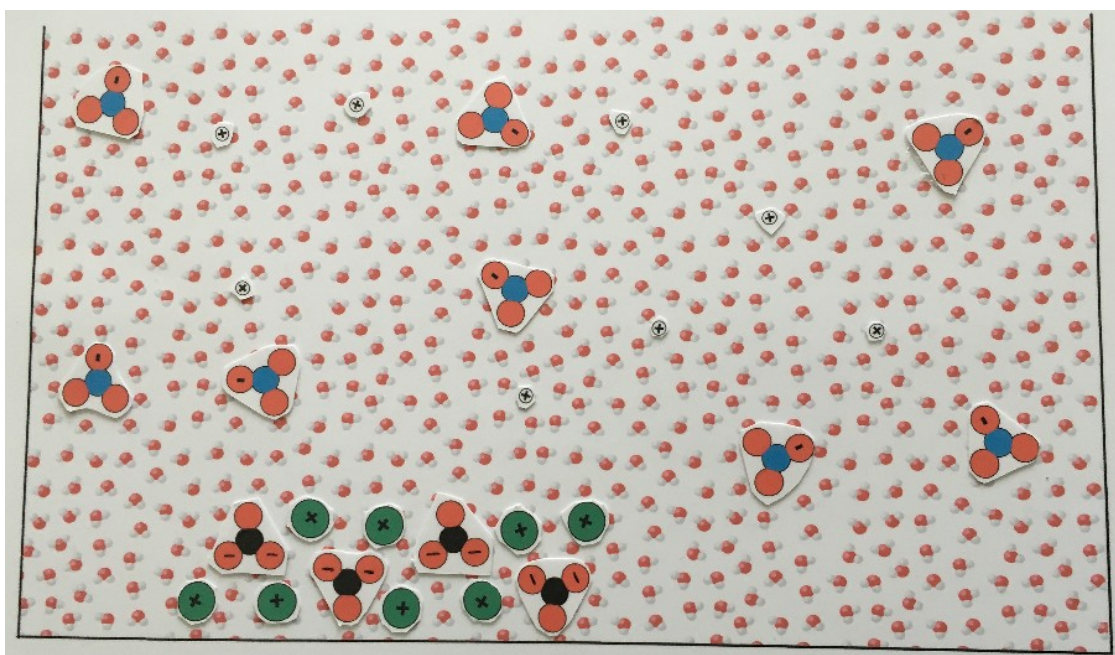
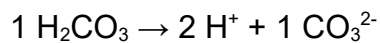
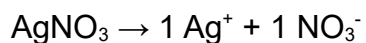
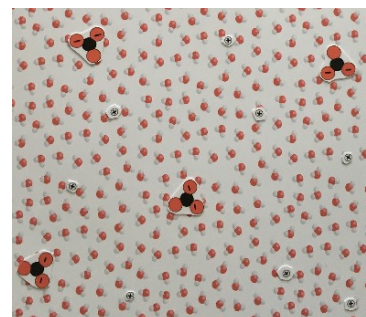
$\text{H}_2\text{CO}_3 (\text{s})^* :$



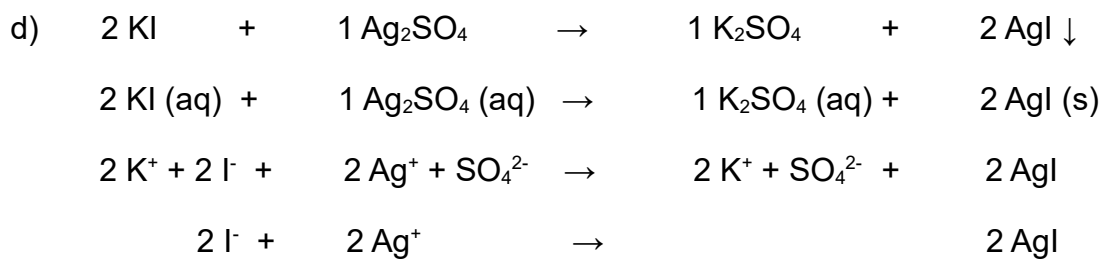
$\text{AgNO}_3 (\text{aq}) :$



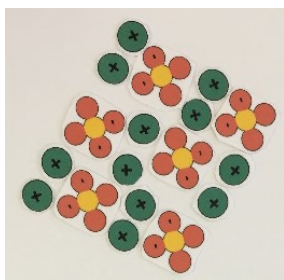
$\text{H}_2\text{CO}_3 (\text{aq}) :$



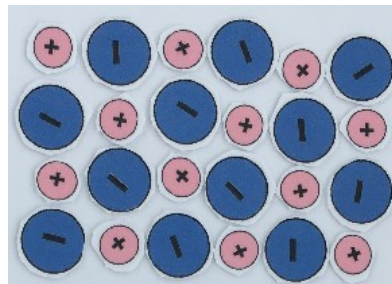
* bien que présenté sous une forme solide pure pour l'exercice, l'acide carbonique ne peut réellement se trouver pur. En effet, on l'obtient par dissolution et réaction de CO_2 avec l'eau et on le trouve donc toujours mélangé à l'eau.



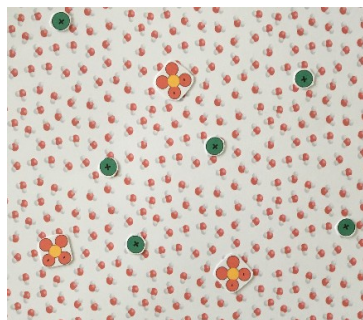
$\text{Ag}_2\text{SO}_4 \text{ (s)} :$



$\text{KI (s)} :$



$\text{Ag}_2\text{SO}_4 \text{ (aq)} :$



$\text{KI (aq)} :$

